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## Preliminary Notes

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### Ring currents in 3:4-benzpyrene, 1:2-benzanthracene, and 1:2;5:6-dibenzanthracene

The work of the PULLMANS and others aimed at relating the electronic configuration of aromatic hydrocarbons to their carcinogenic potency (a theory well summarized by COULSON<sup>1</sup>) has been at least moderately successful. As some of the theory rests on molecular-orbital calculations it is of interest to re-examine the range of validity of molecular-orbital theory as applied to polycyclic hydrocarbons by making theoretical predictions based on the theory which can be checked directly experimentally. In this note, we shall make use of a modification due to MCWEENY<sup>2</sup> of LONDON's<sup>3</sup> theory of diamagnetic anisotropy in aromatic molecules, which is based directly on molecular-orbital theory. We shall obtain theoretical predictions of the  $\pi$ -electron ring currents in several carcinogenic polycyclic hydrocarbons. As these ring currents are responsible not only for diamagnetic anisotropy but also for nuclear-magnetic-resonance chemical shifts the predictions are subject to rather rigorous experimental test. Similar calculations by BERTHIER *et al.*<sup>4</sup>, based on LONDON's unmodified theory, predict the total induced moment, but not the individual ring currents.

MCWEENY has shown that the secondary magnetic field,  $H'$ , at any point, arising from the diamagnetic effect of ring currents in an aromatic hydrocarbon in the presence of an external magnetic field,  $H$ , normal to the plane of the molecule, is given by

$$H' = 2\beta \left( \frac{2\pi e}{hc} \right)^2 \frac{S^2 H}{a^3} (\sigma_1 + \sigma_2),$$

where  $S$  is the area of a benzene ring,  $a$  is the length of a C-C bond,  $\beta$  is the standard resonance integral, and  $\sigma_1$  and  $\sigma_2$  are sums involving bond orders and mutual bond polarizabilities obtained from normal (unperturbed) molecular-orbital theory. The

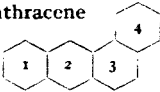
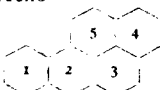
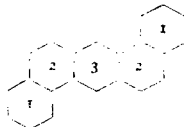
expression can be simplified for practical calculations by a suitable unitary transformation, and it can be shown that the results can be put in the form

$$H' = -2\beta \left( \frac{2\pi e}{hc} \right)^2 \frac{S^2 H}{a^2} \sum_i J_i K(\vec{r}_i),$$

where the sum goes over all rings and the  $K(\vec{r}_i)$  are defined functions. The  $J_i$  are then identified as the ring currents.

We have subjected three polycyclic hydrocarbons to this analysis: the extremely potent carcinogen 3:4-benzpyrene, the mild carcinogenic 1:2; 5:6-dibenzanthracene; and 1:2-benzanthracene, only very weakly active itself, but numbering among its monomethyl and dimethyl derivatives many carcinogens of widely varying potency. The results are given in Table I (compare with the value 1/9 for benzene).

TABLE I  
RING CURRENTS

Molecule	Ring No.	Ring current
1:2-Benzanthracene 	1	0.1243
	2	0.1412
	3	0.0990
	4	0.1247
3:4-Benzpyrene 	1	0.1338
	2	0.1422
	3	0.0933
	4	0.1435
	5	0.1196
1:2; 5:6-Dibenzanthracene 	1	0.1264
	2	0.1048
	3	0.1432

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